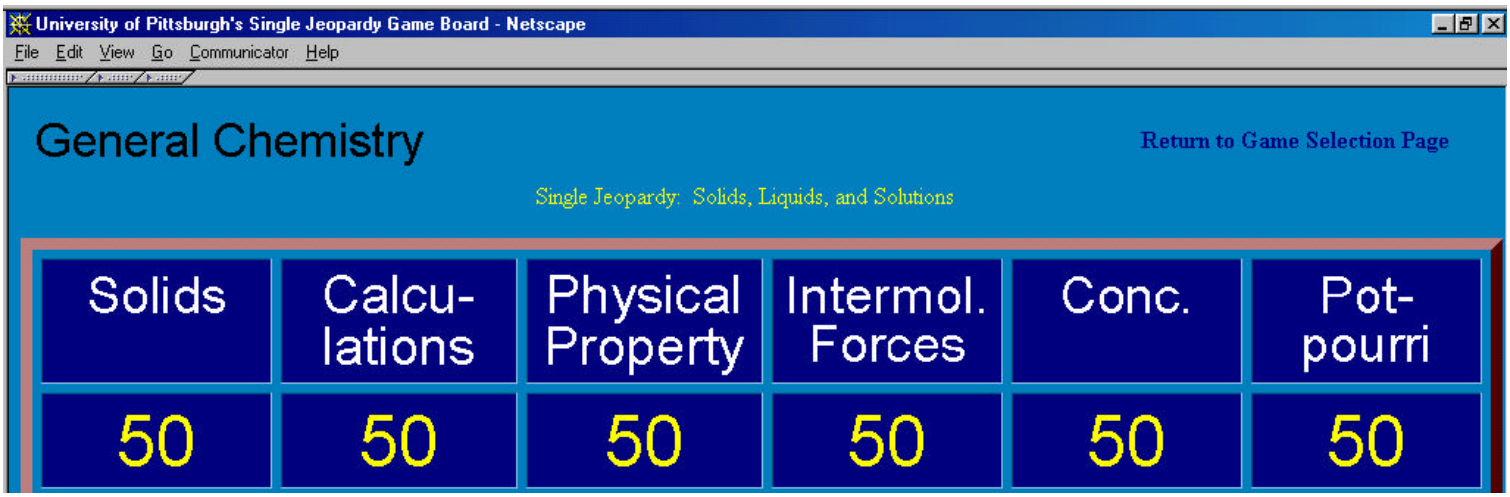

Solids

50	ABAB layers	
100	nearest neighbors in BCC	
150	74.06%	
200	4 atoms per unit cell	
250	also known as FCC	

Calculations

50	k_B if m = .1 and $\Delta T_B = 2$	
100	mol solute in 200 mL of .1M solution	
150	m for .1 mol sugar in 200 g alcohol	
200	$P_A^o = 100$	
250	wt. % for 10 g salt in 100 g H_2O	

Physical Properties

50	$P_{(g)}$ for $H_2O_{(l)} \rightleftharpoons H_2O_{(g)}$	
100	T at vp = 760 Torr	
150	solid $\rightleftharpoons$ liquid $\rightleftharpoons$ gas	
200	3 section plot of P vs T	
250	$P_{CH_3OH} = P^o_{CH_3OH} (1-X_{H_2O})$	

Intermolecular forces

50	among CH_4 molecules	
100	Among NH_3 molecules	
150	$A^+ - B^-$. . . .$A^+ - B^-$	
200	can be up to 40 kJ/mol	
250	among CO_2 in dry ice	

Concentrations

50	$\frac{\text{mol solute}}{\text{V solution (L)}}$	
100	$\frac{\text{mol solute}}{\text{Kg solvent}}$	
150	$(V)_L \times M =$	
200	$\frac{\text{mol A}}{\text{mol A + mol B}}$	
250	$\frac{(\text{mol} \times \text{FW})_A \times 100}{\text{grams total}}$	

Potpourri

50	a unit cell	
100	decreases as T increases	
150	there are 12 of these	
200	2 atoms per unit cell	
250	plot of this gives ΔH_{vap}	