

Solids, Liquids, Solutions

University of Pittsburgh's Single Jeopardy Game Board - Netscape

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General Chemistry

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Single Jeopardy: Solids, Liquids, and Solutions

Solids	Calculations	Physical Property	Intermol. Forces	Conc.	Pot-pourri
50	50	50	50	50	50

Solids

50	ABAB layers	
100	nearest neighbors in BCC	
150	74.06%	
200	4 atoms per unit cell	
250	also known as FCC	

Calculations

50	k_B if $m = .1$ and $\Delta T_B = 2$	
100	mol solute in 200 mL of .1M solution	
150	m for .1 mol sugar in 200 g alcohol	
200	$P_A^\circ = 100$	
250	wt. % for 10 g salt in 100 g H ₂ O	

Physical Properties

50	$P_{(g)}$ for $H_2O_{(l)} \rightleftharpoons H_2O_{(g)}$	
100	T at $v_p = 760$ Torr	
150	solid \rightleftharpoons liquid \rightleftharpoons gas	
200	3 section plot of P vs T	
250	$P_{CH_3OH} = P_{CH_3OH}^\circ (1 - X_{H_2O})$	

Intermolecular forces

50	among CH ₄ molecules	
100	Among NH ₃ molecules	
150	$A^+ - B^- \dots A^+ - B^-$	
200	can be up to 40 kJ/mol	
250	among CO ₂ in dry ice	

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Concentrations

50	$\frac{\text{mol solute}}{V \text{ solution (L)}}$	
100	$\frac{\text{mol solute}}{\text{Kg solvent}}$	
150	$(V)_L \times M =$	
200	$\frac{\text{mol A}}{\text{mol A} + \text{mol B}}$	
250	$\frac{(\text{mol} \times \text{FW})_A \times 100}{\text{grams total}}$	

Potpourri

50	a unit cell	
100	decreases as T increases	
150	there are 12 of these	
200	2 atoms per unit cell	
250	plot of this gives ΔH_{vap}	