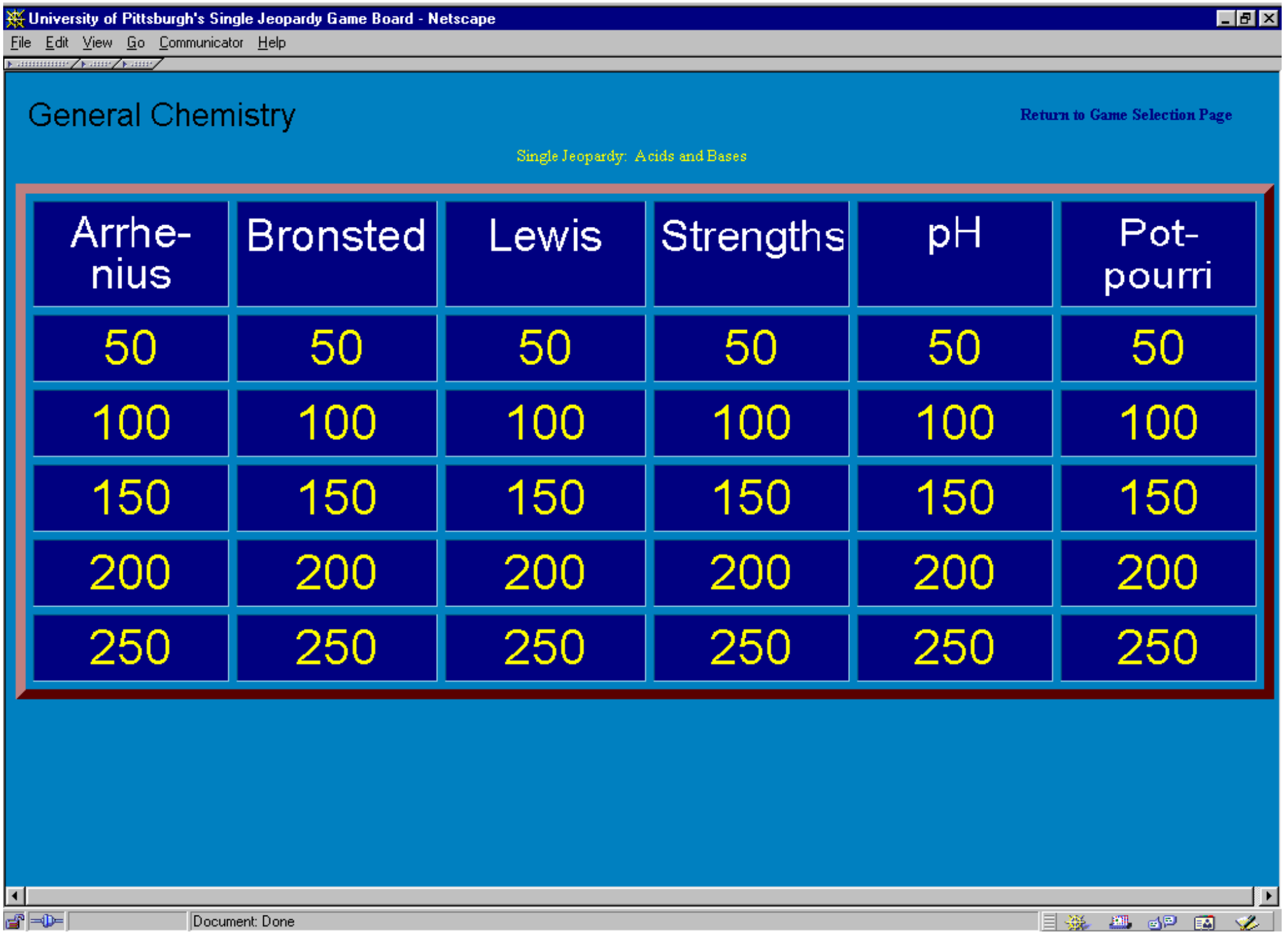
University of Pittsburgh's Single Jeopardy Game Board - Netscape
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General Chemistry
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Single Jeopardy: Acids and Bases
Arrhe-nius
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50 50 50 50 50 50
100 100 100 100 100 100
150 150 150 150 150 150
200 200 200 200 200 200
250 250 250 250 250 250
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Arrenhuis

50	Formula of any base.	KOH
100	formula of any acid	HCl
150	a diprotic acid	H_2SO_4
200	a diprotic base	$Ca(OH)_2$
250	acid + base $\longrightarrow$	salt + water

Bronsted

50	acid	proton donor
100	base	proton acceptor
150	acid reacts to form it	conjugate base
200	not an Arrhenius	NH_3
250	amphoteric liquid	H_2O

Lewis

50	base	electron pair donor
100	acid	electron pair acceptor
150	formula of any acid	BF_3;CO_2
200	formula of any base	NH_3, F^-
250	formed in acid-base reaction	covalent bond

Strengths

50	stronger acid than H_2O	HCl
100	weaker acid than H_2O	OH^-
150	stronger base than H_2O	F^-
200	stronger base than NH_2^-	N^{3-}
250	same as HCl in H_2O	HB

pH

50	halide salt giving pH >7	KF
100	salt giving pH ≃ 7	NaCl
150	any acid anhydride	SO_3
200	example of a buffer	$NH_3(aq)$, $NH_4Cl(aq)$
250	value for normal acid rain	5.6

Potpourri

50	-log[OH]	pOH
100	conjugate acid for CH_3^{-1}	CH_4
150	anhydride of H_3PO_4	P_4O_{10}
200	pH value for 0.1M NaOH	13
250	pH value for .1M $Ca(OH)_2$	13.3